

Download Freezing Point Depression Lab Answers

Molecular Mass by Freezing Point Depression Darlene D'Souza, Nithya Mitta, Ashwini Parchure, Avni Shah
Prelab 2. a.) If the thermometer reading was 1.4 degrees C too high there would be no affect on the molecular mass because the change would be the same. b) If solvent spilled I'm doing a freezing depression lab, and I can't seem to understand what it's asking me! PLEASE, if you have ANY experience with the lab, help me!! So: 1) I found the freezing point of t-butanol (23.01 deg Celcius) 2) I found the freezing point of t-butanol and aspirin (19.76 deg Celcius) SO, FOR THE TRICKY PART: 1 ...Molar mass is useful to identifying an unknown substance; methods such as using finding the freezing point depression of an unknown substance can help in finding the identity. In this laboratory, we will find the molar mass of an unknown substance that was combined with BHT using freezing point depression. The given thermometer should be used to record when the freezing points happen. We find ...Molecular Mass by Freezing Point Depression Kyle Miller November 28, 2006 1
Purpose The purpose of this experiment is to determine the molecular mass of organic compounds which are dissolved in a solvent by noting the depression in the freezing point of the solution as compared to the freezing point of pure solvent since freezing point depends only on the number of particles that are dissolved ...